Primrose Schools
ICSE Curriculum
(A Unit of Primrose Educational Trust, Chennai)
An ISO 9001:2015 Certified Institution

State Purpose of Worksheet – [2018 – 2019]

[KG Class work/Home Assignment/Alternate Class work/Internal Assessment/Continuous Assessment/Unit Test/Cycle Test/Revision Test/Mid Term Examination/Term Examination/Preboard]

Name: **Std: VI**	**Subject: CHEMISTRY**
Date: **Term:III**	**Topic: NA**
Maximum Marks: 60 **Time Duration:**	**Type of Assessment (if internal assessment):NA**

General Instructions

Answers to this Paper must be written on the paper provided separately.
*You will not be allowed to write during the **first 15** minutes.*
This time is to be spent in reading the Question Paper.
The time given at the head of this Paper is the time allowed for writing the answers.

Section I** is compulsory. Attempt **any four** questions from **Section II.
The intended marks for questions or parts of questions are given in brackets [ ].

Section I [30 Marks]

Attempt **all** questions

Q1. a) Write the atomicity of following [5]

i. $CaCO_3$
ii. $MgSO_4$
iii. KCl
iv. NH_4OH
v. O_2

b) Fill in the blanks [5]

i. The ---------- is an example of crystalline separation.
ii. Element and compounds are ----------substances.
iii. A 2% drop in our body's water supply can trigger signs of ----------
iv. Sugar solution is a ---------- mixture.
v. ---------- gas is used as fuel in rocket.

c) State whether true or false [5]

i. The particle which causes pollution is called pollutants.
ii. CNG is a clean fuel.
iii. Pure water freezes at -2˚C and boils at 100˚C.
iv. Naphthalene balls sublimates on heating.
v. Nitrogen is used for manufacture of fertilizers.

d) Match the following: [5]

i.	Incompressible	Diatomic
ii.	Oxygen	Purification step
iii.	Bronze	Nitrogen
iv.	Chlorination	Alloy
v.	CO_2	Lime water milky

e) Differentiate between [4]

i. Respiration and photosynthesis.

ii. Homogenous and heterogeneous mixtures

f) Give example of each types of mixtures: [6]

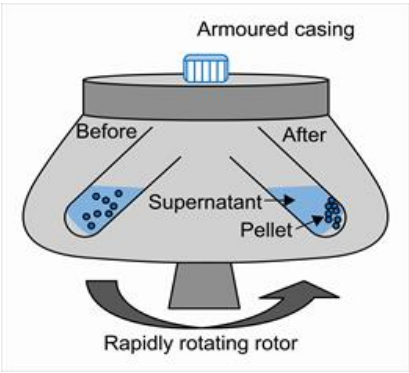
i. Solid –solid mixture
ii. Solid- liquid mixture
iii. Liquid –gas mixture
iv. Gas- gas mixture
v. Solid –gas mixture
vi. Liquid -liquid

Section II [30 Marks]

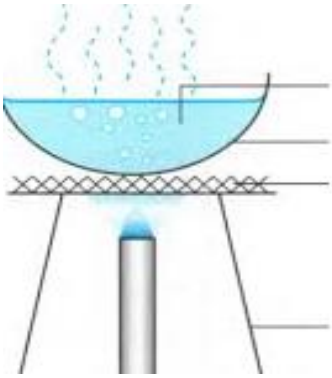
Attempt **any three** questions

Q2 a) Define mixtures and write its four characteristics. [5]

b) Observe the pictures carefully and explain the method of separation with the help of an example. [5]



A



B

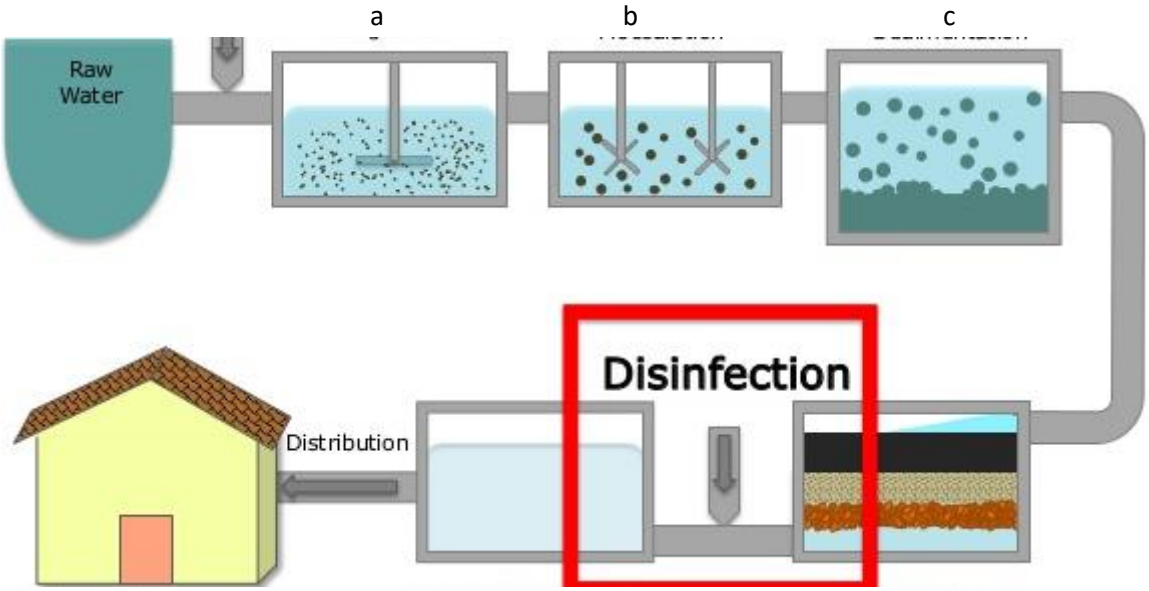
Q3 a) What is potable water. Write its characteristics. [3]

b) Define [2]

i. Saturated solution
ii. Unsaturated solution

c) Label the diagram and Write the process steps a, b, c, d, e.

[5]



Q4. a) Give two example of following: (equation) [4]

i. Water borne disease
ii. Metals and non-metals
iii. Sources of water
iv. Molecules

b) Given here are two diagram of pollution. [6]

i. Explain their causes separately.
ii. What types of pollution is this?

iii. How we can control these types of pollutions.

Q5 a) Find the valency of following elements and write their chemical formulae with hydrogen (atomic no. =1). [6]

C=6, N=7, Cl = 17

b) Draw the orbital diagram for: [4]

i. Potassium(K)=19
ii. Sodium (Na) =11
iii. Neon (Ne)=10
iv. Hydrogen (H) =1