Primrose Schools
ICSE Curriculum
(A Unit of Primrose Educational Trust, Chennai)
An ISO 9001:2015 Certified Institution

State Purpose of Worksheet – II Term Examination [2018 – 2019]

[KG Class work/Home Assignment/Alternate Class work/Internal Assessment/Continuous Assessment/Unit Test/Cycle Test/Revision Test/Mid Term Examination/Term Examination/Preboard]

Name: **Std: VIII**	**Subject: Chemistry**
Date: **Term: II**	**Topic: NA**
Maximum Marks: 80 **Time Duration: 2 hr**	**Type of Assessment (if internal assessment):NA**

Answers to this Paper must be written on the paper provided separately.

*You will not be allowed to write during the **first 15** minutes.*

This time is to be spent in reading the Question Paper.

The time given at the head of this Paper is the time allowed for writing the answers.

Section I is compulsory. Attempt **any four** questions from **Section II.**

The intended marks for questions or parts of questions are given in brackets [ ].

Section I [40 Marks]

Attempt **all** questions

Question 1

a) Fill in the blanks. [5]

i) Fullerenes are ---------- conductor of electricity.

ii) The salts containing water of crystallisation are called ---------- salts.

iii) 10 Å is the particles size in ---------- solution.

iv) The ---------- oxides react with both acids and bases to form salt and water.

v) Chlorine gas has ---------- bonding.

b) Answer in one word. [5]

i) Who discover neutron?

ii) What is the process of formation of metal oxide by heating metal ore in absence of air?

iii) What is heavy hydrogen?

iv) Name one method for removal of temporary hardness from water.

v) Give one example of fast change.

c) Differentiate between the following pairs. [10]
   i) Hard water and soft water
   ii) Acidic oxides and basic oxides
   iii) Reducing agents and oxidising agents
   iv) Colloidal solution and Suspension solution
   v) Exothermic change and Endothermic changes

d) Choose the correct answer from each of the four options given below. [5]
   i) The most reactive metal in following.
      a) Calcium (ca)
      b) Sodium (Na)
      c) Zinc(Zn)
      d) Potassium (K)
   ii) The scientist who discovered proton
      a) Rutherford
      b) Goldstein
      c) Thomson
      d) Crookes
   iii) Water gas is
      a) $(Co+H_2)$
      b) CCl_4
      c) $MgSO_4.\ H_2O$
      d) $CuSO_4.\ 10\ H_2O$
   iv) Which type of reaction produce salt and water as product?
      a) Neutralisation
      b) Combination
      c) displacement
      d) precipitation
   v) What is the valency of Chlorine (Z=17) and calcium (Z=20), where Z is the atomic number.
      a) 3,1
      b) 1,2
      c) 7,2
      d) 2,1

f) Match the items given in the column A with most appropriate ones in column B. [5]

S.no.	**Column A**	**Column B**
i	Acids	Deliquescent
ii	CH_4	Blue litmus to red
iii	Base	Allotrope of carbon
iv	Fullerenes	Red litmus to blue
v	KOH	Hydrocarbon

g) State whether true or false. [4]
   i) Atomic number is the no. of protons in an atom.
   ii) Catalyst can only increase the speed of the reaction.
   iii) $CaSO_4.\ 5H_2O$ is anhydrous salt.
   iv) Precipitate is insoluble substance form in chemical reaction.

h) Write three uses of both graphite and hydrogen. [6]

Section II [40 marks]

*Attempt any **four** questions*

Question 2

a) What is water of crystallisation? Name two salts having no water of crystallisation. [4]

b) Show that, how hydrogen can be prepared by: (Give only the equation) [6]

i) Using an alkali(Na and K)

ii) From reaction between a metal and an acid

iii) Bosch process.

Question 3

a) Balance the following equations: [6]

i) $Zn + H_2SO_4 \rightarrow ZnSO_4 + H_2\uparrow$

ii) $KI + Pb(NO_3)_2 \rightarrow PbI_2\downarrow + KNO_3$

iii) $FeCl_3 + NaOH \rightarrow Fe(OH)_3 + NaCl$

b) Write two properties of both fullerene and graphite. [4]

Question 4

a) Elements A, B, C have atomic number 16, 10 and 18 respectively. [6]

i) Draw their orbital diagram.

ii) Write their valency.

b) Write four uses of hydrogen. [4]

Question 5

a) Define

i) Electrovalent or Ionic bond

ii) Covalent bond [4]

b) Give one example of each of (equations only) the following. [3]

i) Combination reaction

ii) Displacement reaction

iii) Neutralisation reaction

c) What is Greenhouse effect? What are the harmful effect of green house? [3]

Question 6

a) What are Isotopes? Give two example. [1]

b) Define saturated, unsaturated and supersaturated solution. [3]

c) How you can remove temporary hardness and permanent hardness from water. (Give equations) [6]